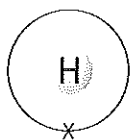
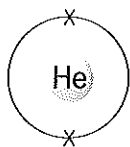


# Electron Arrangement

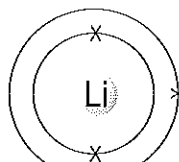
**Q13** Give the full electronic arrangement in the following dot and cross diagrams.  
(The first three have been done for you).



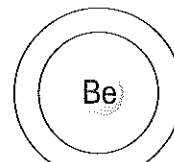
1  
Atomic no. = 1



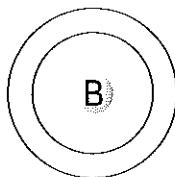
2  
Atomic no. = 2



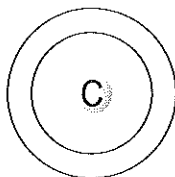
2, 1  
Atomic no. = 3



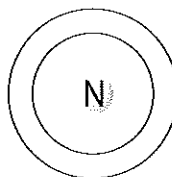
Atomic no. = 4



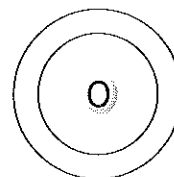
Atomic no. = 5



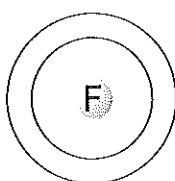
Atomic no. = 6



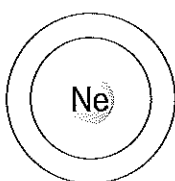
Atomic no. = 7



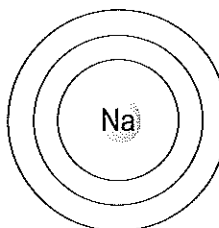
Atomic no. = 8



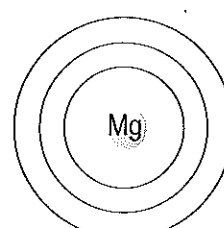
Atomic no. = 9



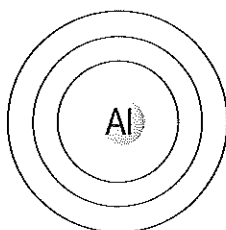
Atomic no. = 10



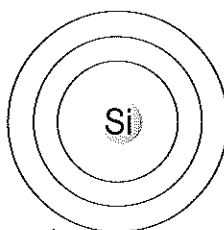
Atomic no. = 11



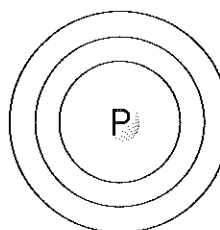
Atomic no. = 12



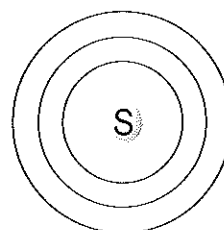
Atomic no. = 13



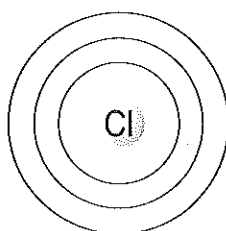
Atomic no. = 14



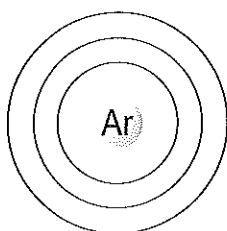
Atomic no. = 15



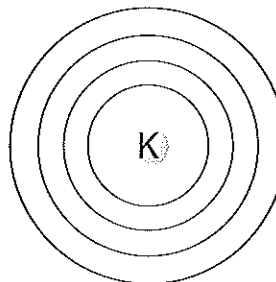
Atomic no. = 16



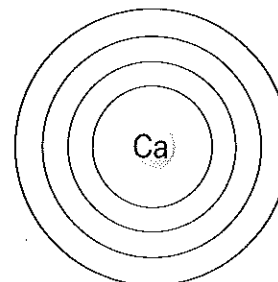
Atomic no. = 17



Atomic no. = 18



Atomic no. = 19



Atomic no. = 20

## Top Tips:

In Exams they're always asking you to draw out electronic arrangements, or "configurations" — just make sure you can work them out from atomic numbers or the Periodic Table. They might only ask you to draw the outer shell — easy or what.....